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TEACHER.CO.KE SERIES 41

233/3
PRACTICAL
PAPER 3

CONFIDENTIAL

The information contained in this paper is to enable the head of the school and the teacher in charge of chemistry to make adequate preparation for this year's mock chemistry practical examination. NO ONE ELSE should have access to this paper or acquire knowledge of its contents. Great care MUSY be taken to ensure that the information herein does not reach the candidates either directly or inderectly. The teacher in charge of chemistry should NOT perform any of the experiments in the same room as the candidates nor make the results of the experiments available to the candidates or give any other information related to the experiments to the candidates related to the experiments to the candidates. Doing so will constituents examination irregularity which punishable.

In addition to the apparatus and fittings found in the laboratory, each candidate will require the following.

1. About 100 cm^3
2. About 60 cm^3 of solution K
3. One burette 0 – 50 ml
4. One pipette 25.0 ml
5. One pipette filter
6. One 250 ml volumetric flask
7. Two 250 ml conical flask
8. One 250 ml beaker (glass)
9. One 100ml measuring cylinder
10. One 25 ml/ 50ml/ measuring cylinder
11. One 10ml measuring cylinder
12. One label
13. Exactly 9.45g of solid C
14. About 20 cm^3 of solution D
15. About 300 cm^3 of 1.0 sulphuric (vi) acid.
16. One thermometer – 10 to 110^0 c
17. One stop watch
18. About 0.5 g of solid N
19. About 0.5g of solid Q

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20. One test – tube holder
21. 6 clean and dry test tubes.
22. One blue, and one red litmus paper
23. One boiling tube
24. About 500ml/ of distilled water in a wash bottle.
25. Bunsen burner
26. Solid sodium hydrogen carbonate
27. pH chart
28. One metallic spatula

ACCESS TO

1. 2M Aqueous Ammonia solution supplied with a dropper.
2. Phenolphthalein indicator supplied with a dropper.
3. Universal indicator solution supplied with a dropper.
4. Solution O, Sodium carbonate solution
5. 2M hydrochloric acid supplied with a dropper.

NOTES

1. Solid C is glucose
2. Solution D is prepared by dissolving 3.16 g of potassium manganet (vii) in about 500 cm^3 of 1M sulphuric acid and diluting to one litre of solution with distilled water.
3. Solution J is 0.1M Hydrochloric acid.
4. Solution K is prepared by mixing 5g og NaOH with1g of $NaNO_3$. Dissolve the mixture in distilled water and dilute to one litre.
5. Solid N is Zn SO_4 $7H_2O$
6. Solid Q is potassium titrate.

NOTE:

The teacher in charge should perform the experiments for question 1 and 2 and draw the table of the results for table 1 and table 2 respectively. The results should be sent together with the students scripts for marking.

Name of teacher who performed the experiments:.....

.....

T S C No: .....

Date: .....

Sign: .....