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CHEMISTRY
PAPER 3
CONFIDENTIAL INSTRUCTIONS TO SCHOOLS

In addition to the fittings and apparatus in a chemistry laboratory, each candidate will require the following

- A burette
- A 25cm^3 pipette
- A pipette filler
- 2 conical flasks
- Distilled water in a wash bottle
- A stand and a clamp
- 150cm^3 of solution A
- 100cm^3 of solution B
- 80cm^3 of solution C
- Boiling tube
- 6 dry test tubes in a rack
- Test tube holder
- Metallic spatula
- 0.5g solid E
- About 2g solid F
- 5mls liquid K

ACCESS TO

- Methyl orange indicator
- 2M sodium hydroxide solution
- Aqueous ammonia (2M $NH_{3(aq)}$)
- 0.5M lead nitrate solution
- Acidified $KMnO_4$
- Acidified $K_2Cr_2O_7$
- Sodium chloride solution

NOTES

1. Solution A is prepared by adding 12.9cm^3 of concentrated hydrochloric acid (specific gravity 1.18) to 600cm^3 of distilled water then top up to one litre
2. Solution B is prepared by dissolving 4g of sodium hydroxide in 600cm^3 of distilled water then top up to one litre.
3. Solution C is prepared by dissolving a mixture of 8.4g of sodium hydrogen carbonate and 1.6g of sodium chloride is about 600cm^3 of distilled water and then top up to one litre.