**233/3 SAMIA SUB-COUNTY 2021 CONFIDENTIAL**

***Apart from the usual laboratory fittings, each student should have the following***:

1. About 0.5g of solid A in a stoppered container
2. About 0.5g of solid B in a stoppered container
3. 100cm3 of solution R
4. 150cm3 of solution Q
5. 150cm3 of solution P
6. Distilled water in a wash bottle
7. About a spatula end –full of solid Calcium Hydroxide
8. Red and blue litimus papers
9. Three 250ml conical flasks
10. One burette0-50ml
11. One pipette 25ml
12. One 50ml measuring cylinder
13. One 10ml measuring cylinder
14. One 25cm3 volumetric flask
15. Phenolphthalein indicator
16. Labels (2)
17. Stop watch
18. Two boiling tubes
19. One metallic spatula
20. Five test tubes on a test-tube rack
21. Wooden splint
22. Test tube holder
23. 100ml glass beaker
24. 1g sodium hydrogen carbonate
25. White paper
26. Filter funnel

***The student should also get access to***:

1. 10% Hydrogen peroxide (freshly prepared + dropper).
2. 2M Barium nitrate solution +dropper
3. 0.5M Hydrochloric acid + dropper
4. Source of heat
5. Acidified potassium manganite (VII)
6. Acidified potassium dichromate (VI)
7. 2M dilute sulphuric (VI) acid
8. Ethanol with a dropper

**NOTES**

* *Solid A is Hydrated ferrous ammonium sulphate*
* *Solution B is Melleic acid*
* *Solution R is prepared by welghing exactly 4.8g of sodium carbonate dissolves it to make 1dm3 of solution*
* *Solution Q is prepared by weighing exactly 172cm3 of hydrochloric acid (35-37% sp.gr 1.18)and dissolving to make 1dm3 of solution*
* *Solution P is prepared by weighing exactly 37.2g of sodium thiosulphste pentahydrate and dissolving to make 1dm3 of solution.*