**NAME …………………………………………………INDEX NO …………………………….**

 **Candidates signature ……………………...**

 **Date ………………………………………**

**233/1**

**CHEMISTRY**

**PAPER 1**

**(THEORY)**

**JULY/AUGUST 2019**

**TIME: 2 ¼ HRS**

**GATUNDU SOUTH JOINT EXAM Kenya Certificate of Secondary Education**

**CHEMISTRY PAPER 1**

**Instructions to candidates**

1. Answer ALL questions in the spaces provided in the question paper
2. Electronic calculators and mathematical tables may be used
3. All working must be clearly shown where necessary
4. Answer the questions in English

**For Examiner’s use only**

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| --- | --- | --- |
| **QUESTION** | **MAXIMUM SCORE** | **CANDIDATES SCORE** |
| 1-29 |  |  |
| **TOTAL SCORE** |  |  |

**Please turn over**

1. A mixture of hexane and water was shaken and left to separate as shown in the diagram below:



State the identity of;

1. **P** ………………………………..……………….

(ii) **W** ………………………………….……………. (2mks)

1. Copper (II) oxide and charcoal are black solids. How would you distinguish between the two solids? (2mks)

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1. Cooking oils comprise of a mixture of compounds which have a boiling point range of 23oC to 27oC.

(i) What evidence is then to support the statement that cooking oil is a mixture? (1mk)

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(ii) Name another experimental technique that could be used to confirm your answer in part **(i)** above. (1mk)

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1. State two uses of hydrogen gas that are also uses of carbon (II) oxide gas. (2mks)

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1. The setup below was used to investigate the reaction between metals and water.

 Identify solid **X** and state its purpose

Solid X ………………..……………………………………………………………….. (1mk)

Purpose ………………………………………………………………………………….. (1mk)

1. (a) Explain why aluminium is a better conductor of electricity than magnesium (2mks)

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(b) Other than cost and ability to conduct, give a reason why aluminium is used for making cables while magnesium is not (1mk)

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1. Differentiate between the bleaching effect of chlorine and sulphur (IV) oxide gases. (2mks)

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1. . (a) The scheme below shows some reactions starting with magnesium oxide. Study it and answer the questions that follow:-
2. Name the reagents used in **steps 2 and 4 (**2mks**)**

Step 2…………………………………………………………………………………

Step 4………………………………………………………………………………….

1. Write an equation for the reaction in **step 2 (**1mk**)**

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1. Describe how a solid sample of anhydrous magnesium carbonate is obtained in **step 5 (**2mks**)**

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1. The formula below represents two cleaning agents M and N.





M: R C OSO3Na+ N: R C COO-Na+

 H H

1. Identify the one that would be suitable to use with water containing calcium ions. Explain. (2mks)

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1. Identify the one that has a longer pollution effect. (1mk)

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1. (a) State Graham’s Law of diffusion. (1mk)

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(b) 240cm3 of oxygen diffused through an orifice in 100 seconds. How long will it take 300cm3 of sulphur (IV) oxide to diffuse through the same orifice? (S = 32, O = 16)(3mks)

1. A hydrated salt has the following composition by mass. Iron 20.2 %, oxygen 23.0%, sulphur 11.5%, water 45.3%. Determine the formula of the hydrated salt (Fe=56, S=32, O=16, H=1). (3mks)
2. When propane is passed over heated broken porcelain, it decomposes into ethane and methane.

(a) What name is given to this type of reaction? (1mk)

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(b) State one application of this reaction. (1mk)

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(c) Name a reagent that can be used to differentiate ethane and methane. (1mk)

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1. (a) Complete the nuclear equation below. (1mk)



 Q R + 

 

 (b) It was found that only 1/32 of radioactive compound  I was remaining after a period of 

 150 days; determine the length of the half-life. (2mks)

1. The diagrams below represent two allotropes of Sulphur. Study them and answer the questions which follow:-



1. Name the **two** allotropes labelled **X** and **Y. (1mk)**

**X ………………………………………………………………………………**

**Y ……………………………………………………………………………..**

(ii) Explain why a piece of burning magnesium continues to burn in a gas jar of Sulphur (IV) Oxide. (2mks)

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1. Describe how you would prepare a dry sample of crystals of potassium sulphate starting with 100cm3 of 1M sulphuric (VI) acid. (3mks)

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1. The solubility of potassium nitrate in water at 70oc is 155g/100g H2O while at 20oc, the solubility is 31g/100g water. 50g of a saturated solution of potassium nitrate at 70oc was cooled to 20oc, calculate the mass that crystallized out. (2mks)
2. Bond energies for some bonds are tabulated below:-

|  |  |
| --- | --- |
| BOND | BOND ENERGY KJ/mol |
| H-H | 436 |
| C=C | 610 |
| C-H | 410 |
| C-C | 345 |

Use the bond energies to estimate the enthalpy for the reaction. (3mks)

C2H4 (g) + H2 (g) C2H6 (g)

1. Nitrogen reacts with hydrogen according to the equation below:-

N2 (g) + 3H2 (g) 2NH3 (g) ΔH = -92KJ

How would the yield of ammonia be affected by increase in:-

 i) Pressure (1mk)

ii) Temperature (1mk)

1. In an electrolysis, a current of 200A was passed through molten oxide of metal **Q** for 58 minutes and 64.8g of the metal deposited. Determine;
2. Charge on metal **Q.** (RMM of Q = 27) (1mk**)**

ii) The volume of oxygen gas produced at standard temperature and pressure

IF = 96500C, molar gas volume s.t.p. =22.4dm3. (2mks)

1. . Consider the reduction potentials below.

Pb2+  + 2e Pb = -O.13V



Mg2+ + 2e Mg = -O.76V

1. Write the overall Redox reaction that takes place when the above half cells are connected. (1mk)
2. Determine the Eθ value of the above cell. (2mks)
3. (a) CFCs have become a big pollution concern this days, what are CFCs. (1mk)

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(b) State two examples of substances that contain CFCs. (1mk)

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(c) State one negative effect of CFCs. (1mk)

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1. The set-up below was used to investigate reaction between copper (II) oxide and ammonia gas 













a) Identify gas T (1mk)

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b) Write an equation for the reaction that took place in the combustion tube. (1mk)

c) State the observation made in the combustion tube. (1mks)

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 23. a) Name the process by which propanol is converted to propanoic acid. (1mk)

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i) What does the experiment show? (1mk) ……………………………………………………………………………………………………………………………………………………………………………………

ii) Name the type of flame shown above (1mk) ……………………………………………………………………………………………………………………………………………………………………………………

iii) Name one characteristic of the flame (1mk) ……………………………………………………………………………………………………………………………………………………………………………………………………

25. a) Sodium chloride dissolves in water to give a neutral solution but aluminium chloride dissolves in water to form Acidic solution. Explain. (2mks)

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 b) Aluminium (III) chloride has a relative formula mass of 267 when in gaseous state. Explain (1mk)

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 26. Write the electronic arrangement of sulphur in the following: (s=16) (3mks)

 i) SO32-

 ii) SO3

 27. a) What is an acid base indicator. (1mk)

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b) Explain why universal indicator may be preferred to acid base indicator. (2mks) …………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………

 28. In the very cold countries, salts are sprinkled on the roads during winter.

 i) Explain why this is important. (1mk) ………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………………

 ii) Give one negative effect of this. (1mk)

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 29. Chlorine gas reacts with cold dilute sodium hydroxide to form a bleaching agent W.

 a) Write the formula of the substance W (1mk)

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 b) Write an equation to show how substance W bleaches. (1mk)

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