

NAME:

SCHOOL:.....

DATE:

SALTS

INSTRUCTIONS TO CANDIDATES

Answer ALL questions in this paper in the spaces provided.

1. Describe how the following reagents can be used to prepare lead (II) sulphate.

Solid potassium sulphate

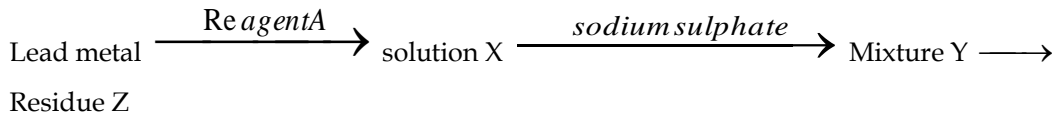
Solid lead (II) carbonate

Dilute nitric acid and distilled water. (3mks)

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2. The reaction below refers to the preparation of lead (II) sulphate starting with lead metal.



Solution

(a) **Name** the type of reaction between solution X and sodium sulphate solution.
(1mark)

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(b) **Write** an ionic equation for the reaction in (a) above (1mark)

(c) **Explain** why it is not possible to prepare residue Z using Lead metal and dilute Sulphuric acid (1mark)

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3. Starting with copper turning, **describe** how a solid sample of Copper (II) Carbonate can be prepared. (3mks)

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4. When lead (II) Carbonate is reacted with dilute sulphuric (VI) acid, the reaction takes place for a short time and then stops. Explain. (2mks)

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