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CHEMISTRY

PAPER 3 (PRACTICAL)

CONFIDENTIAL

The information contained here is to help the teacher in charge of chemistry and head of the institution to prepare adequately before the start of the practical. This information should therefore not be disclosed to anyone else, as that will constitute an exam irregularity.

In addition to the other apparatus found in the laboratory, each candidate should have the following:

1. Complete stand and clamp.
2. 50 ml burette.
3. 25 ml pipette.
4. A white tile.
5. A spatula.
6. Means of heating.
7. 2 boiling tubes.
8. 100 cm³ of solution W.
9. 3.6 g of accurately weighed solid P.
10. A conical flask.
11. 500 ml distilled water in a wash bottle.
12. 100 ml beaker.
13. Means of labeling.
14. 250ml volumetric flask.
15. 6 dry test tubes.
16. About 0.5 g of solid E.
17. One thermometer (-10⁰C to 110⁰C).
18. About 0.5 g of solid F.

Access to each of the following solutions:

NB: Each solution to be supplied with a dropper.

1. 2 M aqueous ammonia.
2. 2 M sodium hydroxide solution.
3. Sodium chloride solution.
4. Barium nitrate solution.
5. Acidified lead (II) nitrate solution.
6. Acidified potassium manganate (VII).
7. Acidified potassium dichromate (VI).
8. Phenolphthalein indicator.

Preparations.

Solid **P** is hydrated oxalic acid.

Solution **W** is **0.2 M** sodium hydroxide solution.

Solid **E** is hydrated aluminium sulphate.

Solid **F** is maleic acid.

