**CATHOLIC DIOCESE OF KAKAMEGA EVALUATION TEST.**

**CHEMISTRY PAPER 2**

**MARKING SCHEME**

a) Noble gases reject rare/inert gases

b) K and W accept Lithium and Potassium

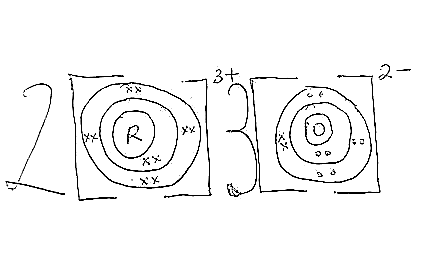
c) Q, hs the fewest number of protons hence experiences weaker nuclear force of attraction.

d) L3M2

e) i) Making electric cables: it is a good conductor of electricity, it is ductile.

ii) Making cooking pans: It is malleable good conductor of heat.

f) R2O3



g) Hassimple molecular structure with weak van der waals forces holding the molecules together.

2.

a) i) Butanoic acid

ii) 3,6-dibromo-5-methylhex-2-ene

iii) ethylpropanoate

b) i) P - Soapless detergent

Q- Soapy detergent.

ii) P, does not form scum/ lathers easily with hard water.

iii) Non- biodegradeable hence pollutes the environment// exppensive

iv) Biodegradeable hence does not pollute the environment// cheap

c) i) ethylbutanoate

ii)butanoic acid& ethanol

3.

a) Decomposition of an electrolyte by passing an electric current through it.

b) -completing the circuit

- maintain electrical neutrality within internal circuit.

c) i) Q=it

= 0.25x7210

=1802.5

ii) 0.9→1802.5

84→?

= 168,233.3333

1 charge→96500 C

? →168,233.333

=1.7434

=+2

d) i) Cu2+(aq) + 2e → Cu(s)

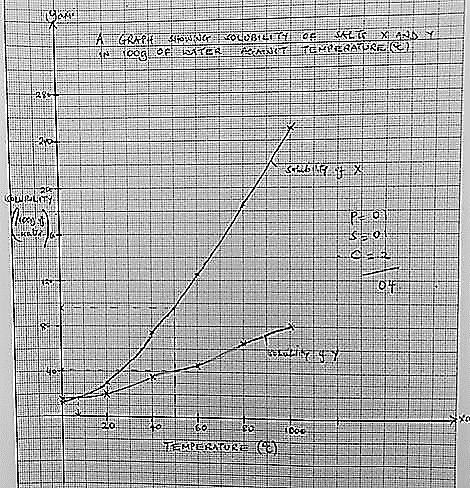
ii) At the beginning, OH-ion are preferentially discharged at the anode where they lose electrons producing O2 gas. After sometime the concentration of OH- decreases and Cl- are preferential discharged due to their higher concentration, they lose electron producing Cl2 gas.

iii) Anode dissolves in the solution, copper anode is an active electrode.

iv) Electroplating to prevent rusting and improve its appearance

4.

a) A solution in which no more solute can dissolve at a particular temperature.



b) i)

ii). X = 96g/100g of water

Y = 41g /100g of water

iii). 7˚C C

c) Hardness that cannot be removed by boiling/ hardness caused by dissolved Mg2+ or Ca2+ ions carnonates.

5.

a) Heat/heating Cuo.

b) Black Cuo turns /changes to reb-brown copper metal.

c) Redox

d) 2CO(g) + O2 (g)→ 2CO2 (g)

e) It is poisonous.

f) i) carbon (II) oxide ii) Copper (II) oxide

g) Ammonia gas // Hydrogen gas (any one)

h) White ash of Na2O remains white// no change ; Na is above C in the reactivity series hence CO cannot reduce Na2O.

6.

a) Concentrated Sulphuric (VI) acid.

b) Denser than air.

c) turns red then eventually bleached to white. Turns red because solution is acidic, turns white because it is bleached

by HOCI

d) Cl2(g) + H20(I)→HOCl(aq) + HCl(aq)

e) PCl5

PCl3

f) Yellow deposits of sulphur; chlorine oxidizes to H2S to S and itself is reduced to hydrogen chloride’

g) Treatment of water (accept any other)

7.a)

i). Downs process

ii) Sodium

iii). Rock salt

iv) Reacts with oxygen , nitrogen and hydrogen carbonates in air forming a coat on the surface.

v) Prevents Cl2(g) from combining with Na(l)

vi) 2Na(s) + 2NH3 (L) 620°C 2NaNH2(L) + H2 (g)

vii) Combined with cyanide used in extraction of gold. etc

b)

i)Bauxite

ii) Bauxite is insoluble in water but soluble in cryolite

iii) O2 produced reacts with carbon anode.

iv) A lot of electricity is used in its extraction.