**Term 1 – 2023 OPENER EXAM**

**CHEMISTRY**

**PAPER 2 (233/2)**

**FORM FOUR (4)**

**Time: 2 Hours**

**Name**: …………………………………………………………. **Adm** **No**: ……………….

**School**: ……………………………………………………….. **Class**: …………………..

**Signature**: …………………………………………………….. **Date**: …………………...

**Instructions;**

1. *Write* ***your name*** *and* ***Admission number*** *in the spaces provided above.*
2. ***Sign*** *and write* ***the date of examination*** *in the spaces provided above.*
3. *Answer all the questions in the spaces provided.*
4. *Mathematical tables and silent electronic calculators many be used.*
5. *All working* ***must*** *be clearly shown where necessary.*
6. ***This paper consists of 14 printed pages. Candidates must check to ensure that no page is missing.***

**FOR EXAMINER’SUSE ONLY**

|  |  |  |
| --- | --- | --- |
| **QUESTION** | **MAXIMUM SCORE** | **SCORE** |
| 1 | 14 |  |
| 2 | 13 |  |
| 3 | 08 |  |
| 4 | 13 |  |
| 5 | 12 |  |
| 6 | 13 |  |
| 7 | 07 |  |
| **Total score** | **80** |  |

***Turn Over***

**1** The diagram below shows part of the periodic table. The letters do not represent the actual symbols of the elements. Study it and answer the questions that follow.

|  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- |
| A |  | | | | | | |  |
|  | B |  |  | C |  |  |  | D |
| E |  | F |  | G |  | H |  |
| I |  |  |  |  |  |  |  |  |

1. Select from the table;

(i) the most reactive metal (1 mark)

…………………………………………………………………………………………

(ii) the least reactive element (1 mark)

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1. State and explain the trend in :

(i) Reactivity of elements **A** and **E**. (2 marks)

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(ii) Atomic radii of elements **F** and **G**. (2 marks)

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1. What name is given to the elements in the same group as element **H**? (1 mark)

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1. Write down the formula of the sulphite of element **E** (1 mark)

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1. The atomic number of an **element J** is 8. Locate **element J** on the grid. (1 mark)

1. Name the type of structure present in :

(i) Chloride of **F** (1 mark)

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(ii) element **C** (1 mark)

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1. A sample of the oxide of **I** was dissolved in distilled water. Both blue and red

litmus papers were dipped into this solution. State and explain the observations made. (2 marks)

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1. Give **one** use of element **B** (1 mark)

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**2** (a) One of the naturally occurring allotropes of carbon is graphite.

(i) Give the other allotrope of carbon. (1 mark)

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(ii) Name **one** use of the allotrope in **a (i)** above (1 mark)

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(iii) Name **one** other element which exhibit allotropy (1 mark)

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(b) Fullerenes are synthetic allotropes of carbon obtained by manipulating carbon using laser beams. Given that the molecular formula of fullerene is C60 , determine its molar mass. (C=12.0) (2 marks)

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(c) Carbon (IV) oxide may be prepared by reacting dilute sulphuric (VI) acid and a carbonate.

(i) Explain why lead (II) carbonate wouldn’t be suitable for use in this reaction.

(2 marks)

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(ii) State and explain the observations made when carbon (IV) oxide is bubbled through a solution of calcium hydroxide. (2 marks)

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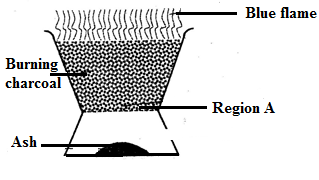
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(iii) Name **one** use of carbon (IV) oxide other than in fire extinguishers.(1 mark)

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(d) The diagram below shows a Jiko when in use. Study it and answer the questions that follow.



**Air**

(i) Write the equation taking place at region **A**. (1 mark)

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(ii) Why is not advisable to place such a Jiko in a closed room? (2 marks)

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**3** (a) What is the effect of impurities on the;

(i) boiling point of water (1 mark)

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(ii) melting point of naphthalene (1 mark)

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(b) Why is rock salt poured on roadways during winter in some countries in Europe? (1 mark)

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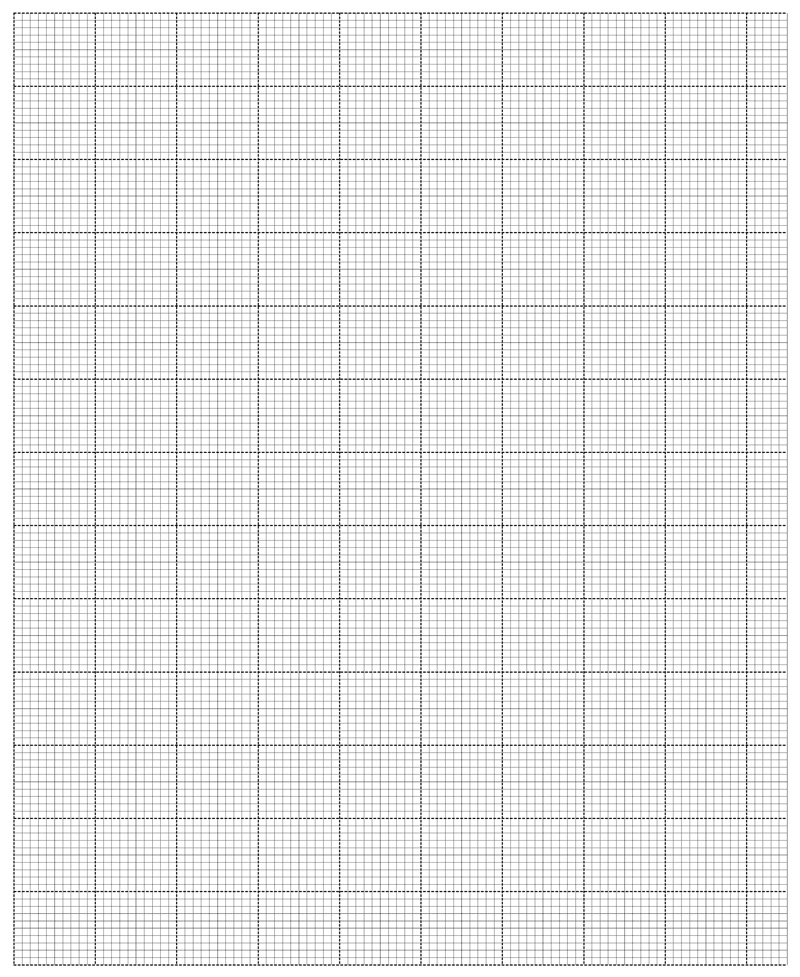
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(c) In an experimentto determine the freezing point of a certain solid, solid **G**, students heat the solid to melting and then allowed to cool. The temperature was recorded after every 30 seconds and the table below was obtained

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| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
| **Time (seconds)** | 0 | 30 | 60 | 90 | 120 | 150 | 180 | 210 | 240 | 270 |
| **Temperature (OC)** | **85.0** | **80.0** | **75.5** | **72.0** | **70.0** | **68.0** | **66.0** | **65.0** | **65.0** | **62.5** |

(i) On the grid provided, plot a graph of temperature (vertical axis) against time.

(3 marks)



(ii) From the graph determine the freezing point of **solid G** (1 mark)

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(d) Give the type of change that occurs when;

(i) Iron nail rusts (½ mark)

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(ii) lead (II) nitrate is heated (½ mark)

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**4** (a) Other than manufacture of dyes and fireworks, Name **two** uses of sulphur (2 marks)

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(b) Sodium sulphite reacts with hydrochloric acid according to the equation below.

Na2SO3(aq) + 2HCl (aq) 2NaCl (aq) + SO2(g) + H2O (l)

25.0 cm3 of 2 M sodium sulphite was reacted with excess hydrochloric acid. Determine the volume of SO2 (g) produced. (Molar gas volume at r.t. p = 24 dm3) (3 marks)

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(c) State and explain the observations made when a moist blue litmus paper is dropped in gas jar full of sulphur (IV) oxide gas. (2 marks)

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(d) Study the flow chart below which shows the preparation of sulphuric (VI) acid and answer the questions that follow

**SO3** (g)

**Air**

**Heat exchanger**

**Catalytic chamber**

**Purifier**

**A**

**Oleum**

**Water**

**SO3** (g)

**Absorption tower**

**Dilution chamber**

**B**

**B**

(i) Give the name of substances;

(I) **A** ……………………………………………………………… (1 mark)

(II) **B** ……………………………………………………………… (1 mark)

(ii) Name the catalyst used in the above process. (1 mark)

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(iii) Write the equation for the reaction between water and oleum (1 mark)

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(iv) State **two** uses of sulphuric (VI) acid (2 marks)

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**5** (a) Give the name of the following organic compounds;

(i) C(CH3)4 (1 mark)

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…………………………………………………………………………………………

(ii) CH3CHCHCH2CH3 (1 mark)

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(iii) CHCCH2CH3 (1 mark)

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(b) (i) Name **two** reagents that can be used to prepare ethyne (1 mark)

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(ii) Write the equation for the reaction that takes place in b (i) above (1 mark)

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(c) Study the flow chart below and answer the questions that follow

**Acidified potassium manganate (VII)**

**B**

**Step III**

**Chloroethane**

**H2, Ni,**

**200 °C**

**Step II**

**A**

**Ethene**

**Step I**

**Ethanol**

**Polyethene**

**Step V**

**Step IV**

(i) Give the reagent and condition for the reactions in **step** **I** and **step III**

**Step I**

Reagent……………………………………..…………………..….. (1 mark)

Condition(s)………………………………………………………... (1 mark)

**Step III**

Reagent……………………………………..…………………..….. (1 mark)

Condition(s)…………………………………………………..…..... (1 mark)

(ii) Draw the structure of compound **B** (1 mark)

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(iii) Give the name of compound **A** (1 mark)

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(iv) State **one** use of Polyethene (1 mark)

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**6** (a) The flow chart below was used in preparation of magnesium hydroxide. Study it and answer the questions that follow.

**Sodium**

**Magnesium**

**Step I**

**Magnesium hydroxide**

**Filtration**

**Products**

**Step III**

**Step II**

**Water**

**Solution L**

**+**

**Gas M**

**Solution K**

**+**

**Gas M**

**Dilute**

**Sulphuric (VI) acid**

(i) State, with a reason the precaution that should be taken in step II (1 mark)

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(ii) Identify;

I solution **K** (1 mark)

…………………………………………………………………………………………

II solution **L** (1 mark)

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(iii) Describe the chemical test for gas **M**  (2 marks)

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(iv) Give **one** use of magnesium hydroxide (1 mark)

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(b) Starting with barium carbonate, describe how a sample containing barium sulphate would be prepared in the laboratory. (3 marks)

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(c) (i) Urea, **(NH2)2CO**, is prepared by passing carbon (IV) oxide through ammonia solution. Explain why urea is a better nitrogenous fertilizer than ammonium nitrate.(C=12.0, N=14.0,H=1.0, O=16.0) (3 marks)

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(ii) Other than manufacture of fertilizer give **one** use of ammonia (1 mark)

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**7** (a) Name the gas produced when concentrated sulphuric (VI) acid is reacted with sodium chloride crystals. (1 mark)

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(b) (i) State **two** uses of chlorine gas. (2 marks)

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(ii) State and explain the observations made when a moist blue litmus paper is dropped in a gas jar full of chlorine. (2 marks)

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(c) DDT is a pesticide which has limited use in disease vector control.

(i) What does DDT stand for? (1 mark)

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(ii) Name **one** environmental impact of DDT (1 mark)

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