**END OF TERM 2 EXAMINATION FORM 4 2022**

**233/3 CHEMISTRY PAPER 3 (PRACTICAL)**

**CONFIDENTIAL**

Each candidate will require.

1. About 100cm3 solutionC1 (H2C2O4.2H2O containing 5.04g in 500cm3 solution)
2. About 100cm3 solution C2 (0.2M NaOH)
3. About 2g solid E (ZnSO4)
4. 2 cm long solid A (Magnesium ribbon)
5. About 60cm3 solution B (0.7M HCl)
6. About 1 g solid F (maleic acid)
7. About 8 cm3 absolute ethanol
8. Burette
9. Pipette
10. Pipette filler
11. 2 boiling tubes
12. 5 test tubes
13. Universal indicator with a chart
14. 0.2g sodium hydrogen carbonate
15. Distilled water
16. 100ml glass beaker
17. Thermometer( -100c to 1100c)
18. Stop watch
19. Means of labelling
20. Spatula

Access to the following

(i).Phenolphthalein indicator

(ii).Pottassium dichromate(vi)

(iii).2M NaOH

(iv).2M Ammonia solution

(v).2M Nitric(v) acid

(vi).Lead nitrate solution

(vii).Barium chloride solution (viii)Bunsen burner