***MECS END OF TERM TWO 2022 CHEMISTRY PAPER ONE MARKING SCHEME PP 1***

1.i*.* substance that produces useful energy when it undergoes a chemical or nuclear reaction. 1MK

ii. -It burns very rapidly producing large amounts of gases which in turn create a huge thrust as they escape.

-It has a very high heat of combustion (4740 kJ mol–1).

-It ignites easily.

ANY 1MK

2. foms positive ions like group I elements½MK

H(g) H+ (g) + e- 1MK

From negative ions like group VII elements½MK

H(g) + e-  H-(g) 1MK

3.i.Rusting

ii. ***–*** Iron nails CHANGES brown./ RUSTS

- Water rises up the delivery tube/water level drops in the trough

4. I. manufacture of glass, soaps and detergents

ii. in paper and textile industries(***accept any***)

b. Ca2+(aq) + CO32-(aq) image CaCO3(s)

Mg2+(aq) + CO32–(aq) image CaCO3(s)

5. -Add water to the solid in a boiling tube and divide the solution into two (1mk)

-add ammonia solution to the two prepared solutions, until in excess (1mk)

A white ppt that ½MK dissolves in excess shows the presence of Zinc ions while a white ppt that doesn’t dissolve shows the presence of aluminium ½MK ions.

6.i. Concentrated sulphuric (VI) acid

ii. Collection over water

drawing and workability

labelling

ConcH2SO4

iii. HCOOH(s) CO(g) + H2O(g)

7. i.a. hydrogen

b. copper(ii) oxide

c. water

ii. it is impure and impurities raise the boiling point

8. i) Metals, the ionic is smaller than the atomic radius

ii. C

9.

|  |  |
| --- | --- |
| **Effect on rate** | **Explanation** |
| increases | Powdered zinc offers a large surface area. |
| increases | particles collide more |

10. a. Rhombic

b. monoclinic

c. Rhombic

11. -Place some flowers in the mortar and crush using the pestle while adding propanone a little at time.Decant the extract in a clean beaker

-Place a drop of the extract at the centre of the filter paper and allow it to spread as far as possible.

-Add a drop of propanone at the centre of the filter paper of the filter paper and allow it to spread as far as possible. A chromatogram showing red and yellow colours is seen.

12.

Q+W=100

Elimination

6Q +7W =694

6Q + 6W= 600

W=94%

Q= 6%

13. i. effervesence/stream of bubbles

ii. White suspension/ precipitate is formed

Ca(s) + H2O (l) Ca(OH)2 (aq) + H2(g)

14. i.**Less volatile**,can displace more volatile acids from their salts

ii.**Oxidising agent**

15. ***Moles of NaOH that reacted***

0.1------------1000

?---------------24.8 =0.00248moles½MK

***Moles of the acid that reacted with the base***

NaOH : HCl

1 : 1

0.00248: 0.00248 moles½MK

***Moles of the acid in 100cm3***

0.2-----------------1000cm3

………………….100cm3 =0.02moles½MK

***Moles of the acid that reacted with the carbonate***

0.02 moles- 0.00248moles =0.01752moles½MK

***Moles of the carbonate that reacted with the acid***

CaCO3 : HCl

1 : 2

? : 0.01752 =0.00876moles½MK

***Mass of the carbonate that reacted***

Moles x molar mass

0.00876 x 100 = 0.876 g½MK

***Percentage of calcium carbonate in limestone***

x100 = 87.6%½MK

16. -sublimes on heating

-It is deliquescent

orange brown, soluble in water

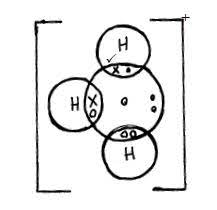
ii- Fe(s) + Cl2(g) FeCl3(s)

17. -Keep substances free from moisture

- Heat substances that require strong heating

- Holding substances being heated

18.



19. a. Burn in a non-luminous flame, carbon burns in a yellow flame, copper (ii) oxide burns in a green flame

b. Add dilute sulphuric (vi) acid, copper (ii) oxide dissolves to form a blue solution while carbon doesn’t dissolve.

c. Heat strongly in a boiling tube, carbon forms a colourless gas that forms a white ppt with while lime water copper(ii) oxide doesn’t.

d. add copper (ii) oxide to each of them and heat strongly, carbon forms a brown substance while copper(ii0 oxide doesn’t.

20.- Add sodium metal into distilled water, to form sodium hydroxide

-Bubble carbon (iv) oxide gas through sodium hydroxide solution for a long period of time to form sodium hydrogen carbonate.

-Evaporate sodium hydrogen carbonate solution to saturate then cool to crystallize

21.i.They are plant extracts that show different colours in acid and base

ii.Do not have a long shelf life

Do not show consistent results

22. -used water softening

-Extraction of some metals

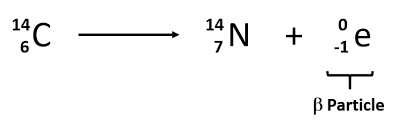
23.IUPAC: International union of pure and applied chemistry

DDT: Dichlorodiphenyltrichloroethane

24.i. Fission: splitting of a heavy nuclide when bombarded with a fast moving neutron

Fusion: small nuclei combine together when made to collide at high velocity (***Both correct***)

ii.



25. 2C + 3H2  +O2 CH3CH2OH

+O2 +O2 +O2

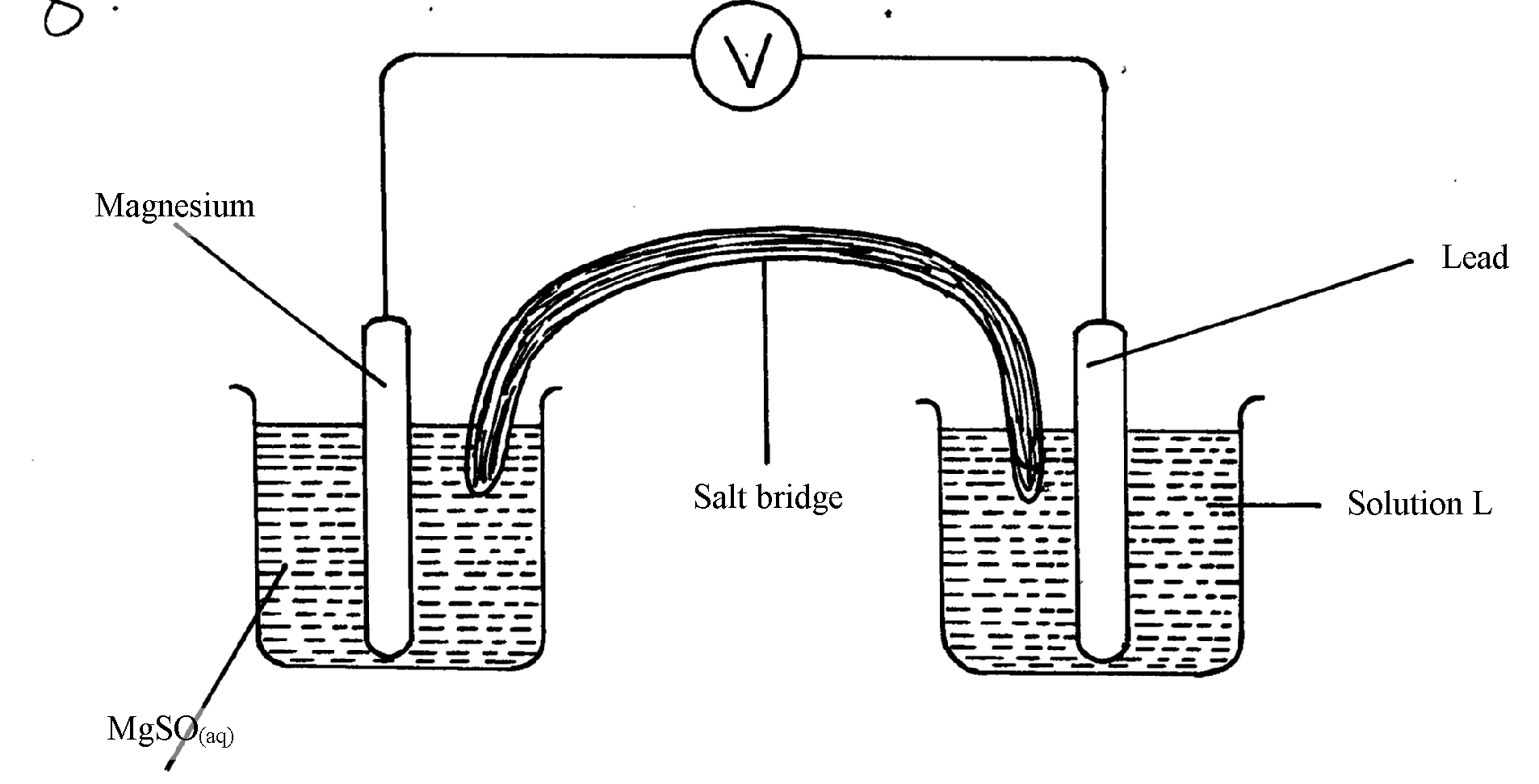
(- 393x2) +(-296x 3)= H + -1386

-1674+ 1386 = -***290kJ/mol***

2 CO2 + 3 H2O

26.i. Pb(NO3)2

ii.



MgSO4(aq)

iii. Mg(s)/Mg2+(aq)//Pb2+(aq)/Pb(s)

27. i.At constant temperature and pressure the rate of diffusion of a gas is inversely proportional to the square root of its density.

ii. 100cm3--------------30sec

150cm3--------------? =45 sec

T2= 2025 x 46/44 = ***46.01 sec***

28. Q-Ethanol

Z- Ethanoic acid

ii. 2CH3CH2OH(l) + 2K(s) 2CH3CH2OK(l) + H2(g)

CH3COOH(aq) + Na2CO3(s) CH3COONa(aq) + CO2(g) + H2O(l)

29i.-The rate of forward reaction is equal to the rate of backward reaction

-at equilibrium the concentration of the reactants and products do not change

ii. pressure will have no effect on the equilibrium because, number of molecule on the reactant and products side are equal.